

Test

Q1-

Sol- NH_3

Q2-

Sol- NH_4NO_3

Nitrogen and hydrogen

Q3-

Sol- (AMU) Atomic mass unit.

Atomic mass is the mass of an atom of particular substance.

Q4-

Sol- H_2O_2

$$1 \times 2 + 16 \times 2$$

$$= 2 + 32$$

$$= \frac{1}{16}$$

Q5-

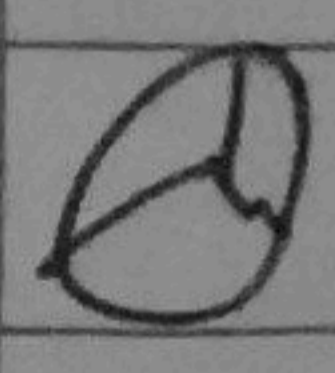
Sol- hg. Mercury.

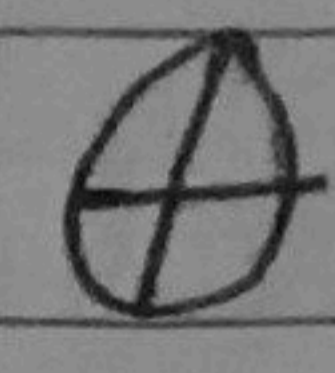
Symbol of Dalton -

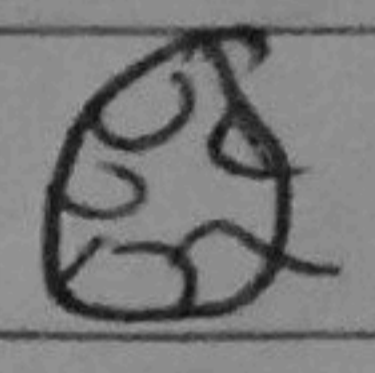
O



O





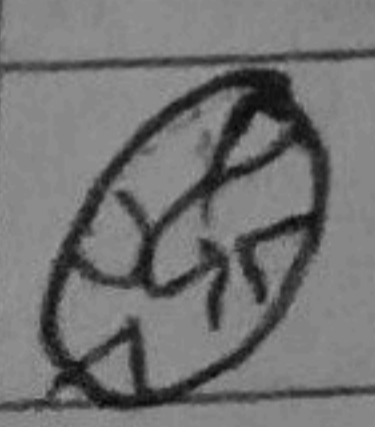


P

L

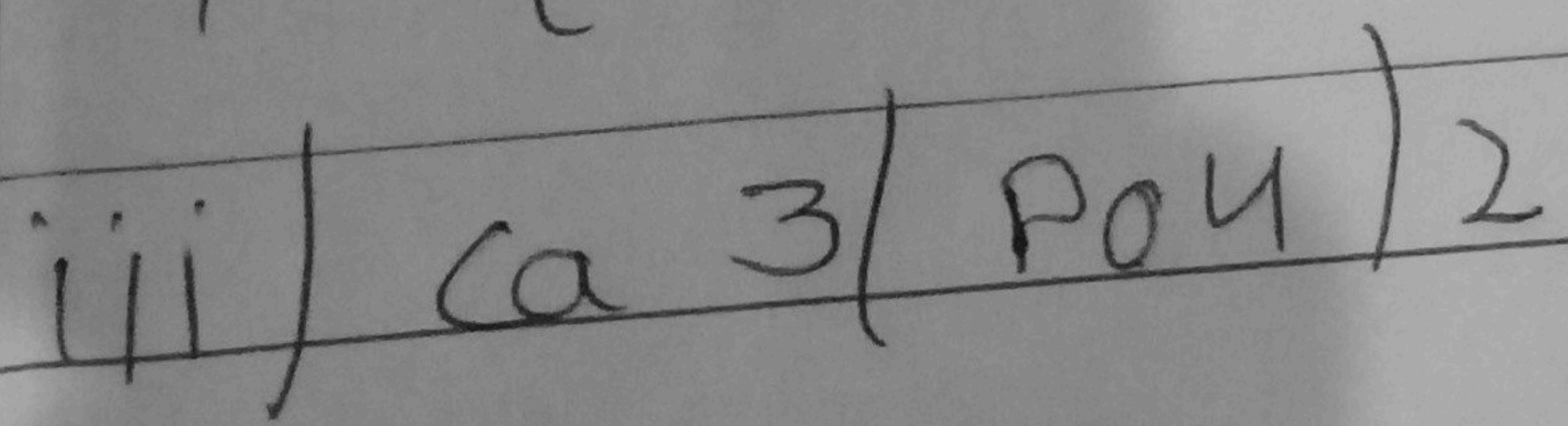
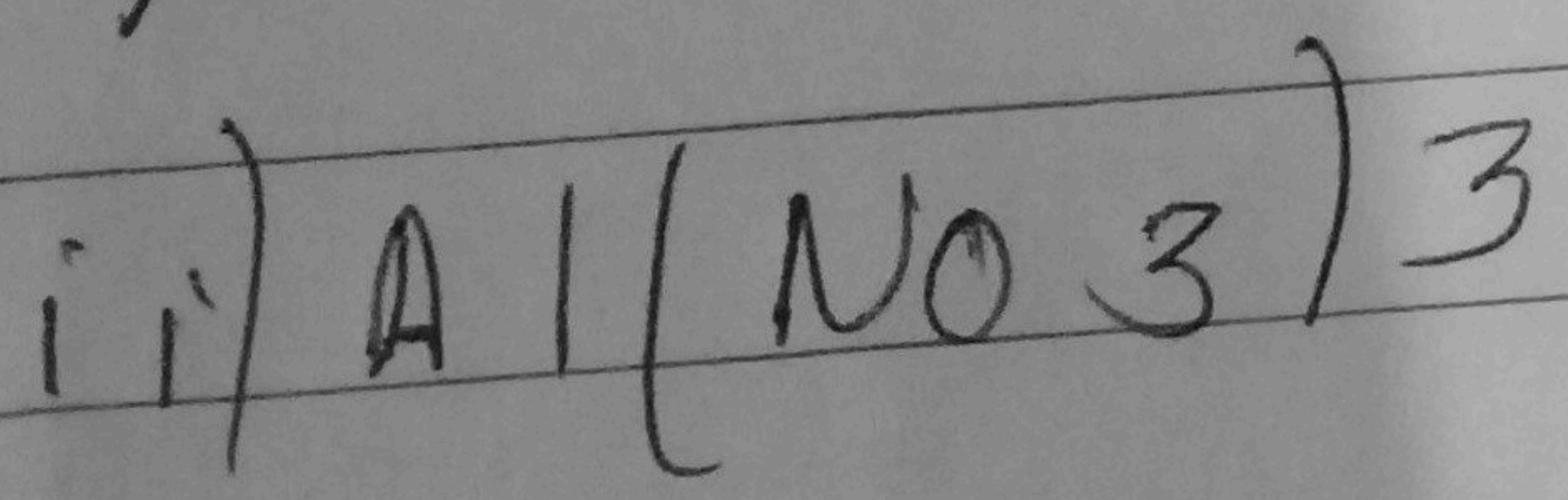
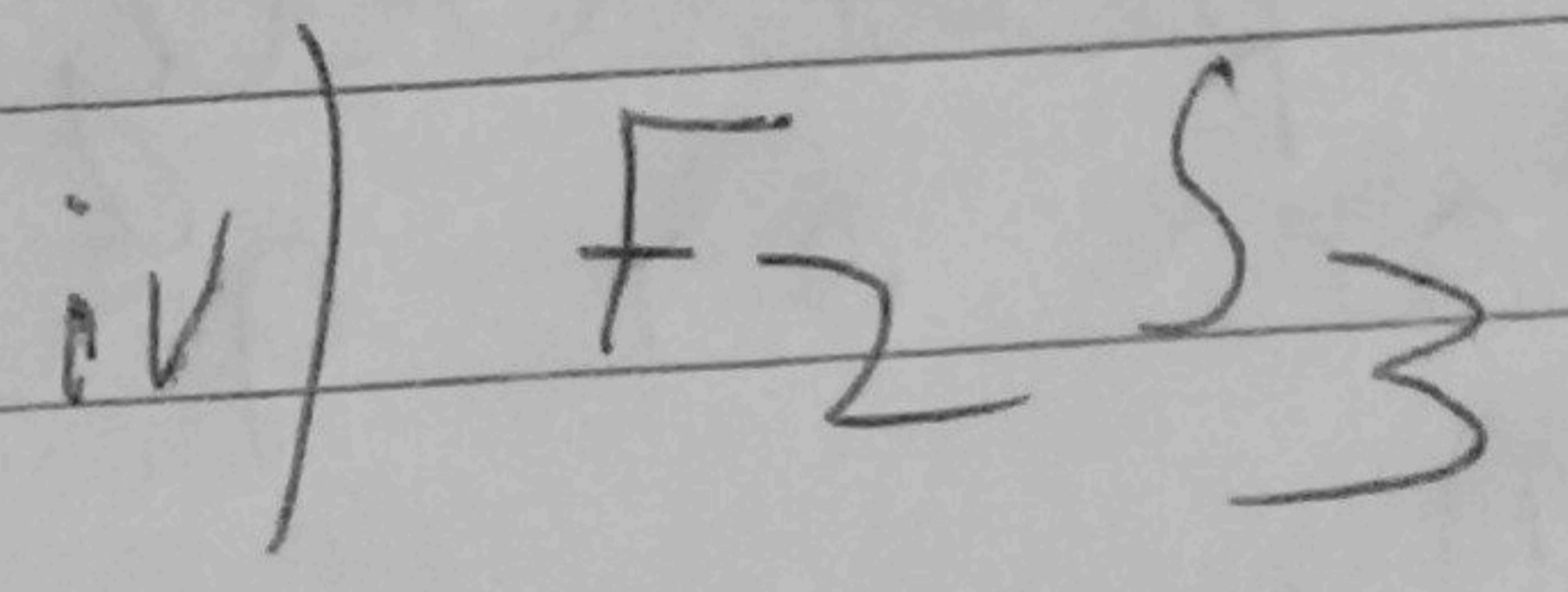
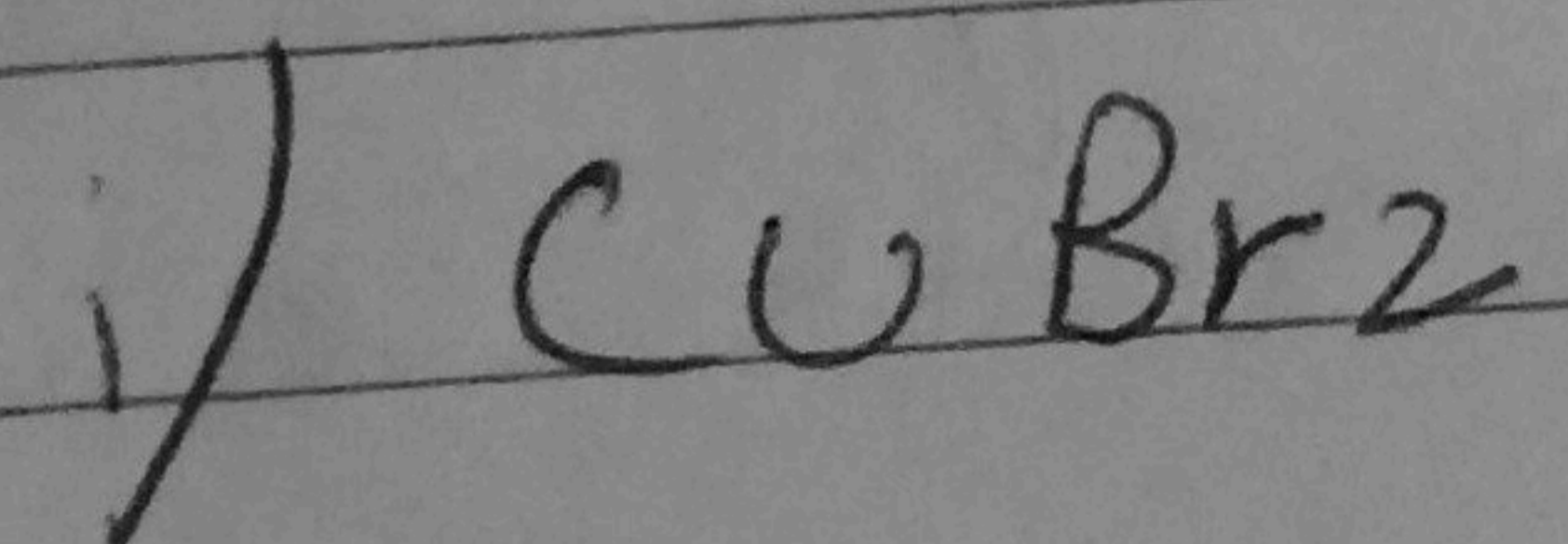
I

C



S

Q 6-
sol-



Q7-
Sol- The number of atoms in molecule of an element.

Helium - 1

Sulphur - 8.

Q8-
Sol- The combining capacity of an atom of element to form chemical bonds is called its valency.
Iron and copper.

- Q9-
sol-
i) CaO
ii) HCl
iii) KOH

10-
Q- A group of two or more atoms of the same or different elements which are chemically bonded are called molecule.

A positively or negatively charged atom or group of atom are called ion.

Q11-

Sol- Copper, Sulphur, Iron, Silicon, SO_2 , PCl_5

Q12-

Sol-

i) All matter contain tiny particles called atom.

ii) Atom of given element is identical by its mass and chemical properties.

iii) ~~AB~~ Atom of different element have different mass and chemical properties.

iv) Atoms are indivisible particles which cannot be created or destroyed in chemical reaction.

v) Atoms combine in the ratio of ^{smallest} whole number to form molecule.

vi) The relative number and kind of atoms are constant in given compound.

ii) 4th postulate

iii) 6th postulate.